

## Moles – practice 1

1 Determine the number of moles present in each of the following:

- (a) 2.3 g of sodium .....
- (b) 0.32 g of O<sub>2</sub> .....
- (c) 1 g of CH<sub>4</sub> .....
- (d) 0.1 g of SO<sub>2</sub> .....
- (e) 4.0 g of N<sub>2</sub> .....
- (f) 2.5 g of Na<sub>2</sub>CO<sub>3</sub> .....
- (g) 15.6 g of Cu(NO<sub>3</sub>)<sub>2</sub> .....
- (h) 2.7 g Fe<sub>2</sub>O<sub>3</sub> .....
- (i) 3 g (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> .....
- (j) 1 g magnesium chloride .....

2 Work out the mass of each of the following:

- (a) 3 mol NaOH .....
- (b) 0.1 mol C<sub>3</sub>H<sub>8</sub> .....
- (c) 0.4 mol CuSO<sub>4</sub> .....
- (d) 100 mol SO<sub>3</sub> .....
- (e) 0.27 mol HNO<sub>3</sub> .....
- (f) 0.85 mol Al<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> .....
- (g) 0.60 mol CaCl<sub>2</sub> .....
- (h) 2.4 mol NH<sub>4</sub>NO<sub>3</sub> .....
- (i) 2 mol calcium carbonate .....
- (j) 0.2 mol sodium hydroxide .....