

## Moles Practice – percentage yield

1 Calculate the theoretical and percentage yield in each case.

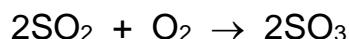
(a) 12.25 g of  $\text{KClO}_3$  react and 6.10 g of  $\text{KCl}$  was formed.



Theoretical yield .....

Percentage yield .....

(b) 0.64 g of  $\text{SO}_2$  reacts with excess oxygen and 0.75 g of  $\text{SO}_3$  is formed.



Theoretical yield .....

Percentage yield .....

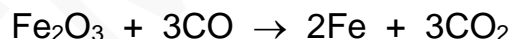
(c) 34 g of  $\text{Sb}_2\text{S}_3$  react with excess  $\text{HCl}$  and 20 g of  $\text{SbCl}_3$  are formed



Theoretical yield .....

Percentage yield .....

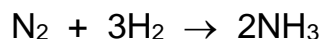
(d) 16 tonnes of  $\text{Fe}_2\text{O}_3$  react with excess carbon monoxide and 10 tonnes of  $\text{Fe}$  are formed



Theoretical yield .....

Percentage yield .....

(e) 200  $\text{cm}^3$  of nitrogen produced 300  $\text{cm}^3$  of ammonia



Theoretical yield .....

Percentage yield .....