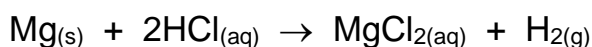


Moles Practice 2

1 Calculate the number of moles in each of the following:

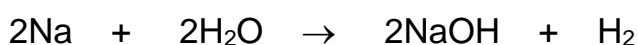
- (a) How many moles of H₂ gas are produced when 0.1 mol of Mg reacts with excess HCl according to the following equation:



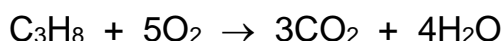
- (b) How many moles of NaCl are produced when 0.5 moles of Na₂CO₃ react with excess hydrochloric acid?



- (c) How many moles of hydrogen gas are produced when 0.4 moles of sodium react with excess water?



- (d) How many moles of O₂ react with 0.01 mol C₃H₈?



- (e) How many moles of H₂S are formed when 2 moles of HCl react with excess Sb₂S₃?



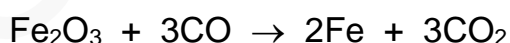
- (f) How many moles of oxygen are formed when 6 moles of KClO₃ react?



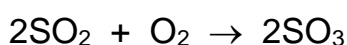
- (g) How many moles of HCl react with 0.2 mol CaCO₃?



- (h) How many moles of iron are formed when 0.9 mol carbon monoxide react with excess iron oxide?



- (i) How many moles of SO₃ are formed when 2.7 mol O₂ react with excess SO₂?



- (j) How many moles of hydrogen would be required to make 2.4 mol NH₃?

