

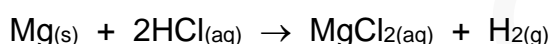
Moles Calculations - practice

1 Determine the number of moles present in each of the following:

- (a) 2.30 g of sodium
- (b) 0.32 g of O₂
- (c) 1.00 g of CH₄
- (d) 12 dm³ of Ne
- (e) 4.00 dm³ of Cl₂
- (f) 100 cm³ of CO₂
- (g) 0.10 g of SO₂
- (h) 400 cm³ of N₂

2 Calculate the number of moles in each of the following:

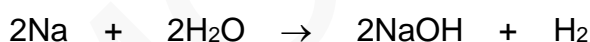
- (a) How many moles of H₂ gas are produced when 0.1 mol of Mg reacts with excess HCl according to the following equation:



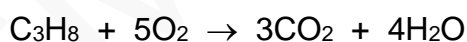
- (b) How many moles of NaCl are produced when 0.5 moles of Na₂CO₃ react with excess hydrochloric acid?



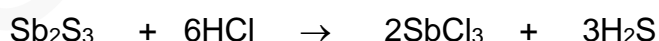
- (c) How many moles of hydrogen gas are produced when 0.4 moles of sodium react with excess water?



- (d) How many moles of O₂ react with 0.01 mol C₃H₈?



- (e) How many moles of H₂S are formed when 2 moles of HCl react with excess Sb₂S₃?



- (f) How many moles of oxygen are formed when 6 moles of KClO₃ react?



- (g) How many moles of HCl react with 0.2 mol CaCO₃?

