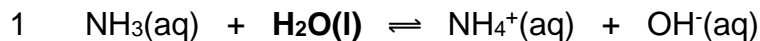


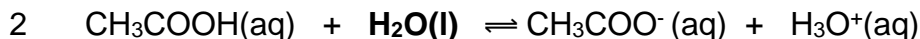
Acids Practice 1

State and explain whether the species in **bold** in each equation is acting as an acid or a base.



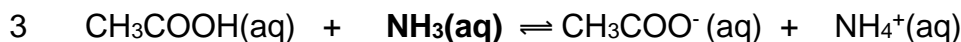
$\text{H}_2\text{O}(\text{l})$ is a(n)..... because.....

.....



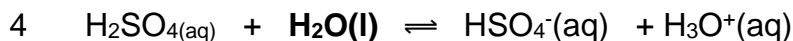
$\text{H}_2\text{O}(\text{l})$ is a(n)..... because.....

.....



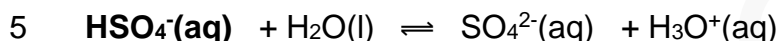
$\text{NH}_3(\text{aq})$ is a(n)..... because.....

.....



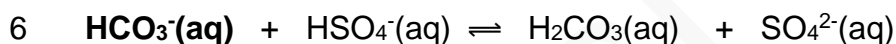
$\text{H}_2\text{O}(\text{l})$ is a(n)..... because.....

.....



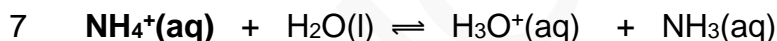
$\text{HSO}_4^-(\text{aq})$ is a(n)..... because.....

.....



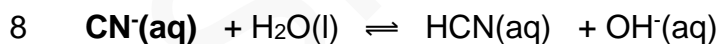
$\text{HCO}_3^-(\text{aq})$ is a(n)..... because.....

.....



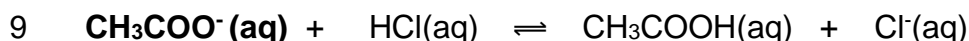
$\text{NH}_4^+(\text{aq})$ is a(n)..... because.....

.....



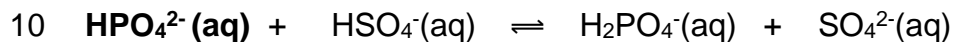
$\text{CN}^-(\text{aq})$ is a(n)..... because.....

.....



$\text{CH}_3\text{COO}^-(\text{aq})$ is a(n)..... because.....

.....



$\text{HPO}_4^{2-}(\text{aq})$ is a(n)..... because.....

.....